

Chemistry 1 a Week Activity		Foundation	8 th Week						
Name:		Class:							
The formula of iron oxide is Fe₂O₃. Calculate the relative formula mass (Mr) of Fe₂O₃. (Relative atomic masses: Fe = 56, O = 16) Relative formula mass = ... Calculate the percentage by mass of iron in iron oxide. (Relative atomic masses: Fe = 56, O = 16) Percentage by mass = ...									
Give the names of the elements combined in barium sulfide. 	Describe the test to show that a gas is carbon dioxide. 								
The formula of the iron oxide is Fe₂O₃. Calculate the maximum mass of iron that can be obtained from 60 tonnes of iron oxide, Fe₂O₃. 									
State the meaning of this hazard symbol.  	Sodium sulfate contains sodium, Na⁺ ions and sulfate, SO₄²⁻ ions. What is the formula of sodium sulfate? 								
What are the electronic configurations of the following ions? <table border="1" style="width: 100%;"> <tbody> <tr> <td style="width: 15%; text-align: center;">K⁺</td> <td>.....</td> </tr> <tr> <td style="text-align: center;">Mg²⁺</td> <td>.....</td> </tr> <tr> <td style="text-align: center;">S²⁻</td> <td>.....</td> </tr> </tbody> </table>		K ⁺		Mg ²⁺		S ²⁻		Calculate the relative formula of Al(NO₃)₃. 	
K ⁺									
Mg ²⁺									
S ²⁻									