

Chemistry 1 a Week Activity		Higher	50 th Week
Name: Class:			
Magnesium displaces zinc from zinc sulfate solution. Complete the ionic equation for the reaction. You should include state symbols. $\text{Mg(s)} + \text{Zn}^{2+}(\text{aq}) \rightarrow \dots\dots\dots + \dots\dots\dots$			
Give the electronic configuration of a potassium ion	Give the formula of a carbonate ion. 		
13.3 cm³ of 0.0500 mol/dm³ citric acid solution was needed to neutralise 25.0 cm³ of sodium hydroxide solution. The equation for the reaction is: $3\text{NaOH} + \text{C}_6\text{H}_8\text{O}_7 \rightarrow \text{C}_6\text{H}_5\text{O}_7\text{Na}_3 + 3\text{H}_2\text{O}$ Calculate the concentration of the sodium hydroxide solution in mol/dm³. Concentration = ... mol/dm³			
HCl is a strong acid. Why is HCl a strong acid? 	Why does increasing the temperature of a reaction increase the rate of reaction? 		
Give the formula of a sulfate ion. 	Give the formula of an ammonium ion. 		
Bromine water is added to hexane and to hexene. What would be observed when bromine water is added to hexane and to hexene? Hexane: Hexene:			