

Chemistry 1 a Week Activity Higher Answers		4 th Week
Name: Class:		
Draw the structure of propene. Show all bonds. $ \begin{array}{c} \text{H} \quad \text{H} \quad \text{H} \\ \quad \quad \\ \text{H}-\text{C}=\text{C}-\text{C}-\text{H} \\ \\ \text{H} \end{array} $	Describe the test to show that a gas is ammonia. Test: moist/damp red litmus. Positive result: if the gas is ammonia the damp red litmus paper will turn blue.	
When decane undergoes complete combustion, a mixture of carbon dioxide and water is formed. Complete the balanced equation for this reaction. $2\text{C}_{10}\text{H}_{22} + 31 \text{O}_2 \rightarrow 20 \text{CO}_2 + 22 \text{H}_2\text{O}$		
In an experiment to produce lead, 7.67 g of lead was obtained. The theoretical yield of lead for the experiment is 8.80 g. Calculate the percentage yield of lead in this experiment. $\text{Percentage yield} = (\text{mass of desired product} \div \text{theoretical yield}) \times 100$ $\text{Percentage yield} = (7.67 \div 8.80) \times 100$ $\text{Percentage yield} = \mathbf{87.2 \%}$		
Describe what is meant by a covalent bond. A bond between two non-metals held together by the attraction between the shared pair of electrons.	State the method used to extract aluminium. Electrolysis	
Name the two substances used to test for the presence of halide ions. 1. Nitric acid. 2. Silver nitrate.	State the formula of a hydrogencarbonate ion. HCO_3^-	
Name the polymer made from propene. Poly(propene)	Give the formula for sodium carbonate. Na_2CO_3	
Give the electronic configuration of a magnesium ion. 2, 8	Give the electronic configuration of a chloride ion. 2, 8, 8	
How do ionic compounds such as sodium chloride conduct electricity? Ionic compounds conduct electricity when <u>molten or aqueous as ions are free to move and carry charge.</u>		