

Chemistry 1 a Week Activity		Higher	48 th Week
Name: Class:			
Calculate the percentage by mass of nitrogen in ammonia (NH₃). Relative atomic masses (Ar): H = 1; N = 14 You must show how you work out your answer. 			
Give the ionic equation for a neutralisation reaction. 		Name the salt produced when ammonia reacts with hydrochloric acid. 	
Balance the following equations. Rb (s) + Li₂SO₄ (aq) → Rb₂SO₄ (aq) + ... Li (s) NO + O₂ → NO₂ C₅H₈O₂ + ... NaH + ... HCl → C₅H₁₂O₂ + ... NaCl			
Ethanoic acid is a weak acid. Why is ethanoic acid considered a weak acid? 		State the functional group for carboxylic acids. State the formula of methanoic acid. 	
Calculate the atom economy for the formation of lithium oxide in this reaction. $\text{Li}_2\text{CO}_3 \rightarrow \text{Li}_2\text{O} + \text{CO}_2$ You must show your working. (relative atomic masses: C = 12, O = 16, Li = 7) (relative formula mass: lithium oxide = 30) Atom economy = ... % Give one reason why reactions with a high atom economy are used in industry. 			