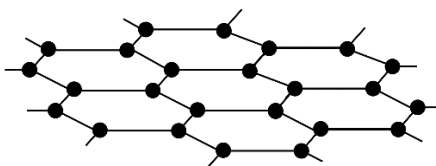


Chemistry 1 a Week Activity		Higher	2 nd Week
Name: Class:			
Name the type of bonding present in Na ₂ CO ₃	State the formula of a sulfate ion.		
Balance the following equation. $\text{Ca(s)} + \text{.....H}_2\text{O(l)} \rightarrow \text{Ca(OH)}_2\text{(s)} + \text{H}_2\text{(g)}$			
Calcium carbonate decomposes on heating to form calcium oxide and carbon dioxide. $\text{CaCO}_3\text{(s)} \rightarrow \text{CaO(s)} + \text{CO}_2\text{(g)}$ 8.000 g of CaCO ₃ was heated strongly for about 10 minutes. 6.213 g of solid remained. Calculate the mass of carbon dioxide gas given off.			
Name the following substance. 	Describe what you would see if damp, blue litmus paper is placed into chlorine gas.		
Describe what is meant by a covalent bond.	Name the following substance. $\begin{array}{ccccc} & \text{H} & \text{H} & \text{H} & \\ & & & & \\ \text{H} & -\text{C} & -\text{C} & -\text{C} & -\text{O}-\text{H} \\ & & & & \\ & \text{H} & \text{H} & \text{H} & \end{array}$		
Calculate the molecular mass of Ca(OH) ₂	Give the formula for a molecule of ammonia.		
Give the electronic configuration for an atom of sodium.	Give the electronic configuration for a sodium ion		
What element is in group 2 and period 4 of the periodic table?			