

Chemistry 1 a Week Activity		Higher	10 th Week				
Name: Class:							
The ionic equation for the reaction between zinc and copper(II) sulfate is: $\text{Zn(s)} + \text{Cu}^{2+}(\text{aq}) \rightarrow \text{Zn}^{2+}(\text{aq}) + \text{Cu(s)}$ Explain why this is described as a redox reaction. 							
Name the two substances that must be present in air for iron to rust. 1. 2.	State the formulae of the following ions. <table border="1"> <thead> <tr> <th>Carbonate</th> <th>Sulfate</th> </tr> </thead> <tbody> <tr> <td>.....</td> <td>.....</td> </tr> </tbody> </table>			Carbonate	Sulfate		
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A student does an experiment to produce a precipitate of lead iodide, PbI₂, using the following reaction. $\text{Pb(NO}_3)_2 + 2\text{KI} \rightarrow \text{PbI}_2 + 2\text{KNO}_3$ He finds that 5.0 cm³ of 0.90 mol/dm³ KI solution reacts with 8.0 cm³ of Pb(NO₃)₂ solution. Calculate the concentration, in mol/dm³, of the Pb(NO₃)₂ solution. 							
Describe a chemical test to show that a liquid contains water. 	What is meant by the term atomic number? 						
Balance the equation for the reaction between sodium hydroxide and sulfuric acid. $\text{.....NaOH} + \text{.....H}_2\text{SO}_4 \rightarrow \text{.....Na}_2\text{SO}_4 + \text{.....H}_2\text{O}$							
What colour flame would potassium ions produce in a flame test. 	State the general formula for alkanes. 						