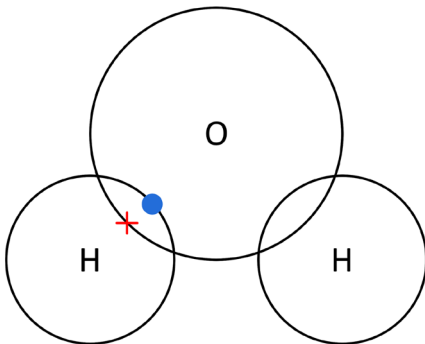


Chemistry 1 a Week Activity		Foundation	10 th Week
Name:		Class: 	
63.0 g of magnalium contains 3.15 g of magnesium. Calculate the percentage by mass of magnesium in the magnalium. 	Water is one of two substances needed for iron to rust. Give the name of the other substance needed for iron to rust. 		
Complete and balance the following equations. $\text{Mg(s)} + \text{.....HCl(aq)} \rightarrow \text{MgCl}_2(\text{.....}) + \text{H}_2(\text{.....})$ $\text{.....ZnO(s)} + \text{C(s)} \rightarrow \text{.....Zn(.....)} + \text{CO}_2(\text{.....})$			
Potassium sulfate contains potassium ions, K^+ , and sulfate ions, SO_4^{2-} . Write the formula of potassium sulfate. 	Aluminium oxide, Al_2O_3 , is made up of ions. Give the number of ions in the formula Al_2O_3		
Give the formula of a molecule of ammonia. 	Give the formula of an ammonium ion. 		
Complete the dot and cross diagram of a molecule of water, H_2O . Show outer shell electrons only. 			
In an experiment, the actual yield of a product is always less than the theoretical yield. Give two reasons why the percentage yield is never 100%. 			
What ion makes a solution acidic? 	What is the electronic configuration of a magnesium 2+ ion. 		