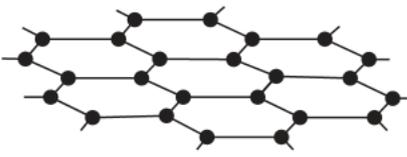


<p>Name:</p>	<p>Class:</p>	<p>2<sup>nd</sup> week</p>
<p>Name the type of bonding present in Na<sub>2</sub>CO<sub>3</sub></p> <p>.....</p>	<p>State the formula of a sulphate ion.</p> <p>.....</p>	
<p>Balance the following equation.</p> $\text{Ca(s)} + \dots\dots\text{H}_2\text{O(l)} \rightarrow \text{Ca(OH)}_2\text{(s)} + \text{H}_2\text{(g)}$		
<p>Calcium carbonate decomposes on heating to form calcium oxide and carbon dioxide.</p> $\text{CaCO}_3\text{(s)} \rightarrow \text{CaO(s)} + \text{CO}_2\text{(g)}$ <p>8.000 g of CaCO<sub>3</sub> was heated strongly for about 10 minutes. 6.213 g of solid remained.</p> <p>Calculate the mass of carbon dioxide gas given off.</p> <p>.....</p> <p>.....</p>		
<p>Name the following substance.</p>  <p>.....</p>	<p>Describe what you would see if damp, blue litmus paper is placed into chlorine gas.</p> <p>.....</p> <p>.....</p> <p>.....</p>	
<p>Describe what is meant by a covalent bond.</p> <p>.....</p> <p>.....</p> <p>.....</p> <p>.....</p>	<p>Name the following substance.</p> $\begin{array}{ccccccc} & \text{H} & \text{H} & \text{H} & & & \\ &   &   &   & & & \\ \text{H} & -\text{C} & -\text{C} & -\text{C} & -\text{O} & -\text{H} & \\ &   &   &   & & & \\ & \text{H} & \text{H} & \text{H} & & & \end{array}$ <p>.....</p>	
<p>Calculate the molecular mass of Ca(OH)<sub>2</sub>.</p> <p>.....</p> <p>.....</p>	<p>Give the formula for a molecule of ammonia.</p> <p>.....</p>	
<p>Give the electronic configuration for an atom of sodium.</p> <p>.....</p>	<p>Give the electronic configuration for a sodium <b>ion</b>.</p> <p>.....</p>	

